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where R is the molar gas law constant. 8.314 J mole⁻¹ K⁻¹, and a 1 is the solvent activity in the solution. This is conventionally taken as unity for the pure solvent, and in ideal solutions, as the mole fraction, x 1. The pressure that must be applied to the solution phase to increase the solvent free energy to that of the pure solvent is the osmotic pressure, π , of the solution.

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er of 10; thus, the 4 in -4.185 is not counted as a significant digit. g. 10⁻⁷.14 = 7.244×10⁻⁸ = 7.2×10⁻⁸ Note that we take the antilog of a number, the digits before the dec-imal point provide information on the power of 10 for the resulting answer; thus, the 7 in -7.14 is not counted as a significant digit. h.

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History of chemistry - Wikipedia
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